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To The Sum Of energy in Form Heat Exchanged Between System And Surroundings (q) And The Work Done On The System (w). $\Delta E_{\text{system}} = Q + W$ But The Laws Of Thermodynamics Are Far More Profound Than Simply Describing The Efficiency Of Steam Engines. T 1th, 2024

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Heat Capacity Of An Object Depends On Both Its Mass And Its Chemical Composition
17 2 Measuring And Expressing Enthalpy Changes In Calorimetry The Heat Released By A System Equals The Heat Absorbed By Its Surroundings Conversely The Heat Absorbed By A System Equals The Heat Released By Its Surroundings, We Will 2th, 2024

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$\text{C}_2\text{H}_5\text{OH}(l) + 3\text{O}_2(g) \rightarrow 2\text{CO}_2(g) + 3\text{H}_2\text{O}(g)$ $\Delta H^\circ_{\text{rxn}} = -1367\text{ kJ/mol}$
A 100.0 mL Sample Of 0.300 M NaOH Is Mixed With A 100.0 mL Sample Of 0.305 M HNO_3 In A Coffee Cup Calorimeter. If Both Solutions Were Initially At 35.0 °C And The Temperature Of The Resulting Solution Was Recorded As 37.0 °C, Determine The $\Delta H^\circ_{\text{rxn}}$ (in Units Of KJ/mol). 4th, 2024

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