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162 CHAPTER 6: THERMOCHEMISTRY To Convert The Answer To Joules, We Write:  $101.3 \text{ J} \cdot 0.18 \text{ L} \cdot \text{Atm} \cdot 1 \text{ L} \cdot \text{Atm}^{-1} = -18 \text{ J}$  6.17 An Expansion Implies An Increase In Volume, Therefore W Must Be  $-325 \text{ J}$  (see The Defining Equation For Pressure-volume Work.) If The System Absorbs Heat, Q Must Be  $+127 \text{ J}$ . The Change In Energy (internal 2th, 2024

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To Thesum Ofenergyin Form Heat Exchanged Between System And Surroundings (q) And The Work Done On The System (w).  $\Delta E_{\text{system}} = Q + W$  But The Laws Of Thermodynamics Are Far More Profound Than Simply Describing The Efficiency Of Steam Engines. T 2th, 2024

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