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Worksheet: Mixed Problems—Mole/Mole Name And Mole/Mass

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Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) 1. $N_2 + 2O_2 \rightarrow N_2O_4$ A. If 15.0g Of N_2O_4 Was Produced, How Many Moles Of O_2 Were Required?
 $15.0g N_2O_4 \times \frac{1 \text{ Mol } N_2O_4}{92.0g N_2O_4} = 0.163 \text{ Mol } N_2O_4$
 $0.163 \text{ Mol } N_2O_4 \times \frac{2 \text{ Mol } O_2}{1 \text{ Mol } N_2O_4} = 0.326 \text{ Mol } O_2$
B. If 4.0×10^{-3} Moles Of Oxygen Reacted, How Many Grams Of N_2 Were Needed?
 $4.0 \times 10^{-3} \text{ Mol } O_2 \times \frac{1 \text{ Mol } N_2}{2 \text{ Mol } O_2} \times 28 \text{ g/mol } N_2 = 0.056 \text{ g } N_2$... Feb 24th, 2024

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