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The Solution Becomes Basic As You Move Past The Neutral Point To PH Extremely Basic Yet The H<sub>3</sub>O<sup>+</sup> Concentration Is Not Zero. It Is Extremely Small Though, 1 X 10<sup>-12</sup> M. At Room Temperature, The Product Of The H<sub>3</sub>O<sup>+</sup> Concentration And The OH<sup>-</sup>concentration In The Solution Will Always Equal 1.0 X 10<sup>-14</sup>. We Can Also Use 3th, 2024

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The  $\text{NH}_4\text{Br}$  is a salt that is soluble in water and so will dissociate into the following ions 100%.  $\text{R NH}_4\text{Br (s)} \rightarrow + \text{NH}_4^+(\text{aq}) + \text{Br}^-(\text{aq})$  | 0.005 0 0 C -0.005 +0.005 +0.005 End 0 0.005 0.005 We are then left with  $\text{NH}_4^+$  ions in solution.  $\text{NH}_4^+$  is a weak acid and thus will reach an equilibrium state. Associate 2th, 2024

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