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Calculations From Chemical Equations Mole - Mole ...7 + 6 KI + 7 H 2SO 4 Cr 2(SO 4) 3 + 4 K 2SO 4 + 3 I 2 + 7 H 2O A) How Many Moles Of Potassium Dichromate (K 2Cr 2O 7) Are Required ... = 407.9 G AgBr This Is The Theoretical Yield 22 B)Calculate The Percent Yield If 375.0 G Of Silver Bromide Was Obtained From The Reaction Theoretical Yield = 407.9 G AgBr Percent Yield = 100 X Actual Yield 1th, 2024Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems)Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) 1. N 2 + 2O 2 → N 2 O 4 A. If 15.0g Of N 2 O 4 Was Produced, How Many Moles Of O 2 Were Required? 15.0g N 2 O 4 1 Mol N 2 O 4 2 Mol O 2 92.0g N 2 O 4 1 Mol N 2 O 4 = 0.326 Mol O 2 B. If 4.0x10-3 Moles Of Oxygen Reacted, How Many Grams Of N 2 Were Needed? 4.0x10-3 Mol O 2 1 Mol N 2 0 4 1 Mol N 2 0 4 2 Mol O 2 92.0g N 2 O 4 1 Mol N 2 O 4 2 Mol N 2 O 4 2 Mol O 2 92.0g N 2 O 4 1 Mol N 2 O 4 2 Mol N 2 O 4 2 Mol O 2 92.0g N 2 O 4 1 Mol N 2 O 4 Mol N 2 O 4 2 Mol O 2 0 2 Nol N 2 O 4 Mol N 2 O 4 2 Mol N 2 O 4 4 Mol N 2 O 4 2 Mol N 2 O 4 Mol N 2 O

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