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4. 200 ML 0.1 M HCl + 100 ML 0.1 M NaOH $\text{HCl(s)} + \text{NaOH(s)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$

This Reaction Goes To Completion Because HCl And NaOH Are Strong Acids And Bases. So, All The NaOH Disassociates Into OH⁻ and Na⁺ Ions And All The HCl Disassociates Into Cl⁻ and H₃O⁺ Ions. There Is A Limiting Reactant In This Problem And It Is NaOH. Calculate The PH 3th, 2024

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