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This Lab, You Will Analyze Energy Transfer Between



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8. On The Back Of This Worksheet, Determine The  
Average Atomic Mass Of Your Element Using This  
Equation: Be Sure To Show Your Work! Average Atomic  
Mass = (avg. Mass Of Isotope)(%) + (avg. Mass Of  
Isotope)(%) + ... (Remember That % Means "pe Mar  
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(4)  $^{238}\text{U}$  ( $t_{1/2} = 4.5 \times 10^9 \text{ Y}$ ) Decays To  $^{234}\text{Th}$  ( $t_{1/2} = 24.1 \text{ D}$ ). Define And Give The Conditions For Secular  
Equilibrium. Assuming The Decay Of  $^{238}\text{U} = 2.3 \text{ Dpm}$   
 $\text{Kg}^{-1}$  What Is The Molar Concentration Of  $^{234}\text{Th}$  At  
Secular Equilibrium. Assume That Here Has Been An  
Extreme Bloom Period With Intensive Scavenging That  
Removes Most Of The  $^{234}\text{Th}$  Jul 2th, 2024

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