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$\Delta G^\circ = -RT \ln K$  Standard Emf Of An Electrochemical Cell:  $E^\circ_{\text{Cell}} = E^\circ_{\text{Ox}} - E^\circ_{\text{Red}} = E^\circ_{\text{Cathode}} - E^\circ_{\text{Anode}}$   
Standard Free Energy Change:  $\Delta G^\circ = -nFE^\circ_{\text{Cell}}$  Where  $F$  Is The Faraday Constant Relationship Of The Standard Emf Of The Cell To The Equilibrium Constant:  
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