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In Your Reaction, The Synthesis Of Banana Oil - Isopentyl Acetate, You Will Be Using Isopentyl Alcohol (3-methyl-1-butanol) And Acetic Acid (ethanoic Acid, See Figure 3). $\text{H}_3\text{C}-\text{O}-\text{O}-\text{H} + \text{HO}(\text{CH}_2)\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3 \rightarrow \text{H}_3\text{C}-\text{O}-\text{O}-\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3 + \text{H}_2\text{O}$ Acetic Acid (ethanoic Acid) MW 60.1, D 1.05 Bp 118 °C Isopentyl Alcohol (3-methyl-1-butanol) MW 88.1, D 0.81 1th, 2024

22. The Fischer Esterification

In This Laboratory Experiment, You Will Prepare Banana Oil (isopentyl Acetate) By The Fischer Esterification Of Acetic Acid And Isopentyl Alcohol (figure 3). The Reaction Will Be Catalyzed By The Addition Of Sulfuric Acid. Once The Reaction Is Complete, A Distillation Will Be Performed To ... 2th, 2024

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1. Esters Of High Volatility - Such As Methyl Formate Methyl Acetate. Ethyl Formate Having Lower Boiling Point Than Those Corresponding Alcohols And Therefore, Can Easily Be Removed From The Reaction Mixture. 2. Esters Of Medium Volatility - These Are Capable Of Removing Water ... 1th, 2024

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