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The Shapes Of The Four Orbitals Are Pictured Below. The S-orbital Is Spherical In Shape. The P-orbital Is A 3-dimensional Dumb-bell Shape. Its 3 Orbitals Exist In Different Planes (called X, Y And Z) And Are Right Angles To One Another. The D-orbital Has Four Lobes, While The F 1th, 2024

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(chapter 2.5 And 2.6) 1. No One Resonance Forms Accurately Depicts The Structure Of The Molecule. The Real Structure Is A Composite Or Hybrid Of All Resonance Forms 2. Resonance Forms Differ Only By The Placement Of  $\pi$ - Or Non-bonding Electrons. Neither The Position Or Hybridization Of The Atoms Changes 2th, 2024

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The First Two Electrs In The System Fill The Orbital 1s And Have The Same Sets Of Four Quantum Numberas As The Two Electrons In ©. The Remaining Electron Must Occupy The Orbital Of The Next Lowest Energy, The 2s Orbital (Figure 3 Or Figure 4). Thus, The Electr 1th, 2024

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