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By Mixing 100.0 G Of Water, H<sub>2</sub>O, And 100.0 G Of Ethanol, C<sub>2</sub>H<sub>5</sub>OH. Determine The  
Mole Fractions Of Each Substance. 2. The Molality Of An Aqueous Solution Of Sugar  
(C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) Is 1.62m. Calculate The Mole Fractions Of Sugar And Water. 3.  
Calculate The Molality Of 25.0 Grams Of 2th, 2024.

K Is The Molal Freezing Point Depression And M Molality Is ...As 1 Atm = 1.013 ×  
10<sup>5</sup> Pa, The Osmotic Pressure Is  $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$  Hence  $C$   
 $= \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$  As 1 M<sup>3</sup> =  
1000 L, This Corresponds To  $C = 1.5 \times 10^{-3} \text{ mol/L}$ , 2024 Richard L. Walker Papers An  
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