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Answers: Moles And Stoichiometry Practice Problems

1) How Many Moles Of Sodium Atoms Correspond To  $1.56 \times 10^{21}$  Atoms Of Sodium?  $1.56 \times 10^{21}$  Atoms Na X  $1 \text{ Mol Na} = 2.59 \times 10^3 \text{ Mol Na}$   $236.022 \times 10 \text{ Atoms Na}$   
2) Determine The Mass In Grams Of Each Of The Following: A.  $1.35 \text{ Mol Of Fe}$   $1.35 \text{ Mol Fe} \times 55.845 \text{ G Fe} = 75.4 \text{ G Fe}$   $1 \text{ Mol Fe}$  B.  $24.5 \text{ Mol O}$  25th, 2024

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SO<sub>3</sub> As A Conversion Factor, We Determine The Mass That This Number Of Moles Of SO<sub>3</sub> Has. The First Step Resembles The Exercises We Did In Section 6.4 "Mole-Mole Relationships In Chemical Reactions". As Usual, We Start With The Quantity We Were Given: 3.59 Mol Fe<sub>2</sub>O<sub>3</sub> × 3 Mol SO<sub>3</sub> / 1 Mol Fe<sub>2</sub>O<sub>3</sub> = 10.77 Mol SO<sub>3</sub> The Mol Fe<sub>2</sub>O<sub>3</sub> Units Cancel, Leaving Mol SO<sub>3</sub> ... 10th, 2024

## **Stoichiometry Problems Worksheet 1 Answers**

Stoichiometry Part 1 Worksheet Answers Stoichiometry Practice Worksheet Solve The Following Stoichiometry Grams-grams Problems: 1) Using The Following Equation: 2 NaOH + H<sub>2</sub>SO<sub>4</sub> → 2 H<sub>2</sub>O + Na<sub>2</sub>SO<sub>4</sub> How Many Grams Of Sodium Sulfate Will Be Formed If You Start With 200.0. 106 People Used View All Course >>. Visit Site. 1th, 2024

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Many Moles Of  $\text{H}_2$  Would Be Required To Produce 5.0  
Moles Of Water? 5.0 Moles Water. B). What Mass Of H  
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